

SRI LANKAN SCHOOL

CLASS : 11

| MONTH | WEEK | D) |
|-----------|------|-------------|
| SEPTEMBER | 1 | 1.16 |
| | 2 | 1.17 |
| | | 1.18 |
| | 3 | 1.19 |
| | 4 | 1.20 |
| OCTOBER | | E) |
| | 5 | 1.23 |
| | 6 | 1.24 |
| | 7 | 1.27 |
| | 8 | 1.23 |
| | 9 | 1.24 |
| NOVEMBER | 10 | 1.25 |
| | 11 | 1.25 & 1.26 |
| | | I) |
| | 12 | 1.48 |
| | 13 | 1.49 |
| | | 1.50 |
| | | 1.51 |
| | 14 | 1.52-1.54 |
| | | 1.55 |
| DECEMBER | | 1.56 & 1.57 |
| | 15 | |
| | 16 | |
| | 17 | |
| | 18 | |
| JANUARY | | B) |
| | 19 | 4.10 |
| | | 4.11 |
| | 20 | 4.12 |
| | 21 | 4.13 |
| | | 4.14 |
| | 22 | 4.15 |
| | | 4.16 |
| 23 | 1.26 | |

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|----------|----|------|
| FEBRUARY | 24 | 1.18 |
| | 25 | 1.18 |
| | 26 | |
| MARCH | 27 | |
| | 28 | |
| | 29 | |
| | 30 | |
| APRIL | 31 | |

| L - MUSCAT | SCHEME OF WORK | AC |
|--|----------------|--------------------|
| SUBJECT: Chemistry | | TEACHER : . |
| RELATIVE FORMULA MASSES(RFM) AND MOLAR VOLUMES OF GASES | | |
| CALCULATE RFM & RAM | | |
| CONCEPT OF MOLES | | |
| INTRODUCING AVOGADRO'S CONSTANT | | |
| MOLE CALCULATIONS I | | |
| MOLAR VOLUME OF GASES AND USAGE | | |
| CHEMICAL FORMULAE AND CHEMICAL EQUATIONS | | |
| DETERMINING CHEMICAL FORMULAE EXPERIMENTALLY | | |
| EMPIRICAL FORMULAE & MOLECULAR FORMULAE | | |
| MOLE CALCULATIONS II | | |
| DETERMINING CHEMICAL FORMULAE EXPERIMENTALLY | | |
| EMPIRICAL FORMULAE & MOLECULAR FORMULAE | | |
| REACTING MASSES | | |
| REACTING MASSES AND PERCENTAGE YIELD | | |
| ELECTROLYSIS | | |
| ELECTRIC CURRENT AND FLOW OF ELECTRONS/IONS | | |
| CONDUCTIVITY OF COVALENT COMPOUNDS | | |
| CONDUCTIVITY OF IONIC COMPOUNDS | | |
| ELECTROLYTES AND NON-ELECTROLYTES | | |
| UNDERSTANDING ELECTROLYSIS AND DEMONSTRATION | | |
| HALF IONIC EQUATIONS | | |
| ELECTROLYSIS CALCULATIONS | | |
| Withdrawal Examination | | |
| Paper Correction and Discussion | | |
| Vacation | | |
| ENERGETICS | | |
| ENDOTHERMIC AND EXOTHERMIC REACTIONS | | |
| CALORIMETRY EXPERIMENTS | | |
| CALCULATE MOLAR ENTHALPY CHANGES | | |
| USE OF "DELTA H" | | |
| ENERGY LEVEL DIAGRAMS | | |
| MAKING AND BREAKING BONDS | | |
| AVERAGE BOND ENERGIES | | |
| Concentration | | |
| Calculations in concentration | | |

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| Introducing titrations |
| Titration calculations |
| Titration calculations |
| Assesment |
| Revision and Past Paper Discussion |
| |
| |
| |
| Mock Examination |

ACADEMIC YEAR 2016 / 2017

JANAKE GUNATHILAKE

| Intended Learning Objectives | Remarks |
|---|---------|
| Calculate the RFM and using RAM | |
| Calculate the number of moles using number of particals and mass | |
| Calculate moles using volumes of gases | |
| | |
| Find the empirical formulae using mass/mass % | |
| Determine the molecular formule using the empirical formulae and molar mass | |
| Find the mass of other substances using mole ration of an equation | |
| Calculate % yield using practical and theoritical data | |
| | |
| Explain method of electrical conductivity in solids gases liquids and aquous solutions | |
| | |
| Calculation of mass of substances deposit of the catode and mass reduction from anode using the amount of current pass through | |
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| | |
| | |
| Calculate the total energy change of a reaction using bond energies and practical data such as temperature change and represent them in energy level diagrams | |
| Find the concentration of asolution given and to calculte | |

